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## Structure Reports

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## Key indicators

Single-crystal X-ray study
$T=100 \mathrm{~K}$
Mean $\sigma(\mathrm{C}-\mathrm{C})=0.002 \AA$
$R$ factor $=0.054$
$w R$ factor $=0.147$
Data-to-parameter ratio $=18.5$
For details of how these key indicators were automatically derived from the article, see http://journals.iucr.org/e.
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## N -9H-Fluoren-9-ylidene-3,4-dimethylaniline

The title compound, $\mathrm{C}_{21} \mathrm{H}_{17} \mathrm{~N}$, was synthesized by the $p$-toluenesulfonic acid-assisted Schiff base reaction between 9-fluorenone and 3,4-dimethylaniline. The crystal structure has been determined at 100 K .

## Comment

Our interest in imines results from the potential use of chiral amines as auxiliaries in the separation of racemic mixtures of aldehydes and ketones. Chiral amines can form Schiff bases with racemic aldehydes and ketones to form diastereomeric mixtures of imines. These diastereomers can potentially be resolved by either flash chromatography or differential recrystallization.

(I)

To test the ease with which imines can be formed, a series of imines were made by reacting various aromatic amines with 9 -fluorenone. A previous study has shown that an aromatic amine with an electron-donating group forms the imine with 9 -fluorenone readily and in good yield (Glagovich et al., 2004). The title compound, (I), was made in order to test what effects disubstitution of the amines will have on imine formation.

The title compound crystallized with molecules in general positions. The planes defined by the dimethylphenyl and the fluorenylidene moieties are nearly at right angles to each other [89.61 (3) ${ }^{\circ}$ ]. An ORTEP-3 (Farrugia, 1997) diagram of (I) is shown in Fig. 1.

## Experimental

In a 100 ml round-bottomed flask equipped with a Hickman still and reflux condenser, 2.5 g ( 14 mmol ) 9-fluorenone, $2.04 \mathrm{~g}(16.8 \mathrm{mmol})$ 3,4-dimethylaniline, 0.014 g ( 74 mmol ) $p$-toluenesulfonic acid, and 50 ml toluene were combined. The resulting mixture was heated under reflux for 24 h . After this time, the orange solution was concentrated under reduced pressure to produce an orange solid. The solid was purified by flash chromatography ( $\mathrm{SiO}_{2}, 80 \%$ hexanes-ethyl acetate) which produced 3.36 g of (I) as an orange solid ( $87 \%$ ). $R_{F}$ $0.59\left(\mathrm{SiO}_{2}, 80 \%\right.$ hexanes-ethyl acetate); m.p. 355.8 K ; IR $\left(\mathrm{CHCl}_{3}\right)$ 3040, 3061, 3017, 1649, 1450, 1376, 1304, 790, 730, $650 \mathrm{~cm}^{-1} ;{ }^{1} \mathrm{H}$ NMR $\left(300 \mathrm{MHz}, \mathrm{CDCl}_{3}\right): \delta 7.92(d, 1 \mathrm{H}, J=7.4 \mathrm{~Hz}), 7.60(d, 2 \mathrm{H}, J=7.4 \mathrm{~Hz})$, $7.46(t, 1 \mathrm{H}, J=7.4 \mathrm{~Hz}), 7.35(d, 1 \mathrm{H}, J=7.8 \mathrm{~Hz}), 7.32(d, 1 \mathrm{H}, J=$ $7.5 \mathrm{~Hz}), 7.16(d, 1 \mathrm{H}, J=7.8 \mathrm{~Hz}), 6.95(t, 1 \mathrm{H}, J=7.5 \mathrm{~Hz}), 6.81(s, 1 \mathrm{H})$,

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$6.73(t, 1 \mathrm{H}, J=7.8 \mathrm{~Hz}), 2.33(s, 3 \mathrm{H}), 2.29(s, 3 \mathrm{H}) ;{ }^{13} \mathrm{C}$ NMR ( 300 MHz , $\left.\mathrm{CDCl}_{3}\right): \delta 162.65,149.42,143.73,141.75,137.62,137.48,132.05,131.64$, $131.55,131.26,130.30,128.33,127.59,127.09,123.21,120.27,120.09$, $119.50,115.50,19.85,19.29$; UV/Vis $\left(\mathrm{CH}_{2} \mathrm{Cl}_{2} ; \lambda_{\text {max }}, \log \varepsilon\right) 404 \mathrm{~nm}$, 2,685; MS calculated for $\mathrm{C}_{21} \mathrm{H}_{17} \mathrm{~N}: \mathrm{M}^{+}: 283$, measured: 283. Analysis calculated for $\mathrm{C}_{21} \mathrm{H}_{17} \mathrm{~N}$ : C 89.01, H 6.05, N $4.94 \%$; found: C 89.26, H 6.18, N 4.81 .

Crystal data
$\mathrm{C}_{21} \mathrm{H}_{17} \mathrm{~N}$
$M_{r}=283.36$
Triclinic. $P \overline{1}$
$a=8.47(3) \AA$
$b=8.870(3) \AA$
$c=10.376(4) \AA$
$\alpha=78.886(6)^{\circ}$
$\beta=79.375(6)^{\circ}$
$\gamma=89.483(6)^{\circ}$
$V=752.1(5) \AA^{\circ}$

$$
\begin{aligned}
& Z=2 \\
& D_{x}=1.251 \mathrm{Mg} \mathrm{~m}^{-3} \\
& \text { Mo } K \alpha \text { radiation } \\
& \text { Cell parameters from } 6350 \\
& \quad \text { reflections } \\
& \theta=2.3-28.4^{\circ} \\
& \mu=0.07 \mathrm{~mm}^{-1} \\
& T=100(2) \mathrm{K} \\
& \text { Parallelepiped, orange } \\
& 0.40 \times 0.34 \times 0.10 \mathrm{~mm}
\end{aligned}
$$

## Data collection

## Bruker SMART APEX

diffractometer

## $\omega$ scans

Absorption correction: multi-scan (SADABS; Sheldrick, 2003) $T_{\text {min }}=0.967, T_{\text {max }}=0.995$
7401 measured reflections

## Refinement

Refinement on $F^{2}$
$R\left[F^{2}>2 \sigma\left(F^{2}\right)\right]=0.054$
$w R\left(F^{2}\right)=0.147$
$S=1.03$
3722 reflections
201 parameters
H -atom parameters constrained
All H atoms were placed in calculated positions, with $\mathrm{C}-\mathrm{H}$ distances of 0.95 and $0.98 \AA$ (for aromatic and methyl H atoms, respectively), and were included in the refinement in the ridingmodel approximation, with $U_{\text {iso }}(\mathrm{H})=1.2 U_{\text {eq }}$ (carrier atom).

Data collection: SMART (Bruker, 1997-1999); cell refinement: SMART; data reduction: SAINT-Plus (Bruker, 1997-1999); program(s) used to solve structure: SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997);


Figure 1
A view of (I) (Farrugia, 1997). Displacement ellipsoids are drawn at the $50 \%$ probability level. H atoms have been omitted for clarity.
molecular graphics: ORTEP-3 (Farrugia, 1997); software used to prepare material for publication: SHELXL97.

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